

06/22/21

Write the electron configuration for the Sulfur atom. How many electrons in sulfur are in p orbitals? And is sulfur paramagnetic?



10e⁻ in p orbitals



Yes paramagnetic

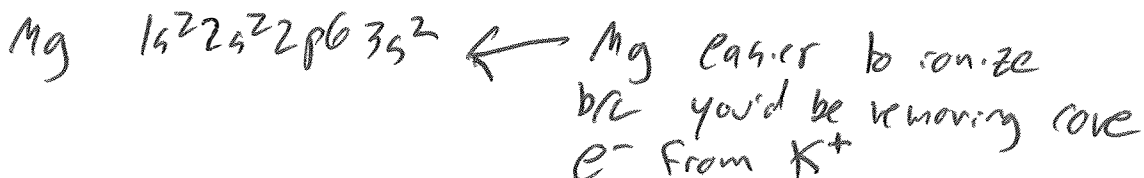
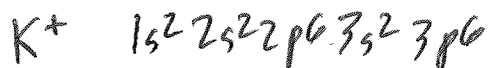
Write the electron configuration of the aluminum ion. Is the aluminum ion paramagnetic?

Al³⁺ has 10e⁻



Not paramagnetic

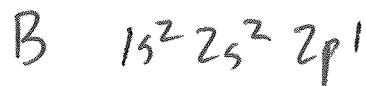
Which is easier to ionize, the potassium ion or the magnesium atom? Explain



Which is larger, the rubidium (Rb) atom or the cesium (Cs) ion? Explain.

Rb atom < Cs atom but Cs⁺ ion is much smaller
than Cs atom so Rb atom > Cs⁺ ion

Which has the larger electron affinity, the boron atom (B) or the beryllium atom (Be)? Explain.



← to add e⁻ to Be you are adding initial e⁻
to 2p, in B, there is already an e⁻
in 2p

