

July 9th recap

Colligative Properties

Raoult's Law $P = X_{\text{solvent}} P_{\text{solvent}}$

boiling point elevation $\Delta T_b = k_b m$

freezing point depression $\Delta T_f = k_f m$

osmotic pressure $\pi = MRT$

* given all these equations on exams

* constants also given on exam

- determine the vapor pressure when a solute (electrolyte and/or non-electrolyte) is added to solvent
- determine boiling point or freezing point when a non-electrolyte or non-electrolyte is added

- determine the molar mass
of a compound using
freezing pt depression equation,
boiling pt elevation equation
or osmotic pressure equation

- determine osmotic pressure
of a solution